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Chapter 19 Acids Bases And

An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous

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solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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STUDY. PLAY. Acids-taste sour-

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conductors of electricity-can turn litmus paper a different color (red)-reacts with certain metals-contains more hydrogen ions than hydroxide ions-reacts with carbonates. Bases-taste bitter-conductors of electricity

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Chapter 19 - Acids & Bases Chapter 19 covers properties of acids and bases, the different definitions, $[H^+]$ and $[OH^-]$ and pH and pOH, K_a and K_b and their use when dealing with weak acids and bases and buffers.

**Chapter 19 - Acids & Bases - Mrs.
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produce.... what did geologist find
interesting abo.... -taste sour...

-conductors of electricity... -can turn
litmus paper.... -taste bitter...

-conductors of electricity... -feels

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slippery... -can....

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2012. Terms in this set (43) Acids taste

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sour, will change the color of an acid-base indicator, and can be. strong or weak electrolytes in aqueous solution.

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The Arrhenius definition states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+ ions). Bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solutions.

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Acids and Bases. SC6. Obtain, evaluate, and communicate information about the properties that describe solutions and the nature of acids and bases. f. Use mathematics and computational thinking

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to compare, contrast, and evaluate the nature of acids and bases in terms of percent dissociation, hydronium ion concentration, and pH.

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Chapter 19 Acids, Bases, and Salts Acids and Bases: Basics These are really just a specific type of chemical compound.

Chapter 19 Acids, Bases, and Salts

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Answer C, a conjugate base is what is
left of an ion after it has donated a
proton in solution.

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SECTION 19.1 ACID-BASE

THEORIES(pages 587-593) This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.

Chapter 19: Acids and Bases

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Homework Packet (50 pts)

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acids taste sour, will change the color of
an acid-base indicator, and can be
strong or weak electrolytes in aqueous
solution.

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