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Stoichiometry
Workbook
Answers

Stoichiometry Workbook Answers

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Stoichiometry Workbook Answers

How many sodium atoms are needed to react with 1.25×10^{24} molecules of water?

Ans: 1.25×10^{24}
atoms Na 1 mol H₂O
2 mol Na 6.02×10^{23}

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atoms Na 1.25×10^{23}
molecules H_2O $\times \times \times$
 $= 6.02 \times 10^{23}$
molecules H_2O 2 mol
 H_2O 1 mol Na 24

Stoichiometry Worksheet #1

continued 5. Hematite,
 Fe_2O_3 , is an important
ore of iron.

Stoichiometry Worksheet 1 Answer Key - Studylib

Answer: 8.75 g O_2 (1
mol O_2 32.00 g O_2) (2
mol H_2 1 mol O_2) (2.02

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g H₂ (1 mol H₂) = 1.10

g H₂ (In your

calculator: $8.75 \div$

$32.00 \times 2 \times 2.02 =$)

13.3 Mass-Volume

Stoichiometry OR Molar

Mass gas @ STP Recall:

Avogadro's Molar

Volume is 22.4 L/mol

for a gas only at STP

Steps: 1) If given

grams, use MM as your

conversion factor to

get to moles of the

given

Chapter 13

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Stoichiometry

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Chapter 03 - Stoichiometry

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the atomic weights for
the atoms in a
chemical formula • So,
the formula weight of
calcium chloride, CaCl_2 ,
would be Ca: $1(40.1$
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choose the correct
answer and follow the

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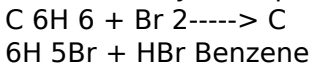
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Stoichiometry (Worksheet) - Chemistry LibreTexts

Stoichiometry Example



(C_6H_6) reacts with Bromine to produce bromobenzene ($\text{C}_6\text{H}_5\text{Br}$) and hydrobromic acid. If 30. g of benzene reacts with 65 g of bromine and produces 56.7 g of bromobenzene, what is the percent yield of the

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Workbook
reaction? 30.g 65 g
56.7 g

Answers

Chapter 3 Stoichiometry - Chemistry

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Mr. Christopherson / Stoichiometry

Ch 12 Stoichiometry

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b. Chapter 12.1, 12.2

Stoichiometry p1

Stoichiometry Basic

Introduction, Mole to

Mole, Grams to Grams,

Mole Ratio Practice

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Problems This chemistry video tutorial provides a basic introduction into stoichiometry.

Ch 12 Stoichiometry Workbook Answers Prentice Hall

Create a new spreadsheet in your workbook and name it 'Practice 1'. In this spreadsheet, create a table using formulas in Excel that can calculate the grams of

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Al (s) and Cl₂ (g)
required to produce a
desired amount of
AlCl₃ (s) according to
the reaction $2 \text{Al}(s) + 3\text{Cl}_2 (g) \rightarrow 2 \text{AlCl}_3 (s)$
Set up your table in a
similar manner to that
shown in Worked
Example #1.

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ecf8427e.

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